

**Show All Work To Receive Credit!** Conversion factors and prefixes:

G = 10<sup>9</sup>, M = 10<sup>6</sup>, k = 10<sup>3</sup>, c = 10<sup>-2</sup>, m = 10<sup>-3</sup>, μ = 10<sup>-6</sup>, n = 10<sup>-9</sup>, p = 10<sup>-12</sup>, 2.54 cm = 1 in, 12 in = 1 ft, 5280 ft = 1 mile, 3 feet = 1 yd, 60 sec = 1 min, 1 hr = 60 min, 4 quarts = 1 gal, 2 pints = 1 quart

1. (6 Pts) Perform each of the following conversions. You must show the complete setup.

a. Convert 88 mg to pg. 
$$\frac{88 \cancel{\text{mg}}}{\cancel{\text{mg}}} \times \frac{10^{-3}}{10^{-12}} = 88 \times 10^9 \text{ or } 8.8 \times 10^{10} \text{ pg}$$

b. Convert 85 μL to nL. 
$$\frac{85 \cancel{\mu\text{L}}}{\cancel{\mu}} \times \frac{10^{-6}}{10^{-9}} = 85 \times 10^3 \text{ or } 8.5 \times 10^4 \text{ nL}$$

2. (6 Pts) Assume each of following numbers are measurements. Perform the indicated operations and then report the answer with the proper number of significant figures.

a. 12.145 cm + 15.1265 cm + 25.2 cm = 
$$\begin{array}{c} \text{To this place} \\ \downarrow \\ 52.47 \text{ or } 52.5 \text{ cm} \end{array}$$

b. 10.25 cm x 12.10 cm x 10.145 cm = 
$$\begin{array}{c} 4 \text{ sig Figs} \\ 1258 \text{ cm}^3 \end{array}$$

c. (25.2 + 103) ÷ (12 x 1.03) = 
$$\begin{array}{c} \text{1st to this place gives } 128.2 \\ \downarrow \\ 10.4 \text{ or } 10 \\ \leftarrow \text{then 2 sig Figs} \end{array}$$
 (assume these all measurements)

3. (4 Pts) A poster measures 22 cm by 44 cm. Determine its area in square inches (inches<sup>2</sup>) (you may ignore significant figures).

$$\frac{22 \cancel{\text{cm}}}{2.54 \cancel{\text{cm}}} \times \frac{44 \cancel{\text{cm}}}{2.54 \cancel{\text{cm}}} \times \frac{1 \text{ in}}{2.54 \cancel{\text{cm}}} \approx 150 \text{ in}^2$$

4. (4 Pts) How many milli-inches are in 7 kilo-feet (You may ignore significant figures)?

$$\frac{7 \times 10^3 \cancel{\text{ft}}}{1 \cancel{\text{ft}}} \times \frac{12 \cancel{\text{in}}}{10^{-3}} = 84,000,000 \text{ or } 8.4 \times 10^7 \text{ m in}$$

5. (5 Pts) A sample of silver ore was found to contain 0.55 % silver by mass. How many mg of silver can be recovered 900.0 kg of ore?

$$\frac{900 \cancel{\text{kg ore}}}{1 \cancel{\text{kg}}} \times \frac{10^3}{10^{-3}} \times \frac{\text{m}}{100 \text{ ore}} \times 0.55 \text{ Ag} = 495,000 \text{ or } 4.95 \times 10^5 \text{ mg Ag}$$

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 12 in = 1 ft, 5280 ft = 1 mile, 3 feet = 1 yd, 60 sec = 1 min, 1 hr = 60 min, 4 quarts = 1 gal, 2 pints = 1 quart

1. (6 Pts) Perform each of the following conversions. You must show the complete setup.

a. Convert 327 pL to mL. 
$$\frac{327 \cancel{\text{pL}} | 10^{-12} | \cancel{\text{m}}}{\cancel{\text{p}} | 10^{-3}} = 327 \times 10^{-9} \text{ or } 3.27 \times 10^{-7} \text{ pL}$$

b. Convert 805 μg to cg. 
$$\frac{805 \cancel{\mu\text{g}} | 10^{-6} | \cancel{\text{g}}}{\cancel{\mu} | 10^{-2}} = 805 \times 10^{-4} \text{ or } 8.05 \times 10^{-2} \text{ cg}$$

2. (6 Pts) Assume each of following numbers are measurements. Perform the indicated operations and then report the answer with the proper number of significant figures.

a. 13.1 cm + 12.526 cm + 0.052 cm = 25.678 or 25.7 cm  
 To this place  
 ←  
 3 sig. figs.

b. 1.212 cm x 6.12 cm x 12.145 cm = 90.0848 or 90.1 cm<sup>3</sup>  
 ↓  
 1st to this place gives 158.2

c. (55.2 + 103) ÷ (13 x 1.03) = 11.8 or 12 (assume these all measurements)  
 ↑  
 Then 2 sig figs

3. (4 Pts) A poster measures 33 cm by 45 cm. Determine its area in square inches (inches<sup>2</sup>). (You may ignore significant figures)

$$\frac{33 \cancel{\text{cm}} | 1 \text{ in}}{2.54 \cancel{\text{cm}}} \times \frac{45 \cancel{\text{cm}} | 1 \text{ in}}{2.54 \cancel{\text{cm}}} = 230 \text{ in}^2$$

4. (4 Pts) How many milli-inches are in 0.5 kilo-feet (You may ignore significant figures)?

$$\frac{0.5 \cancel{\text{kft}} | 10^3 | 12 \text{ in} | \text{mm}}{\cancel{\text{k}} | 1 \text{ ft} | 10^{-3}} = 6,000,000 \text{ or } 6 \times 10^6 \text{ mm in}$$

5. (5 Pts) A sample of silver ore was found to contain 0.75 % silver by mass. How many mg of silver can be recovered 500.0 Mg of ore?

$$\frac{500.0 \text{ Mg ore} | 10^6 | 0.75 \text{ Ag} | \text{mm}}{\cancel{\text{M}} | 100 \text{ ore} | 10^{-3}} = 3,750,000,000 \text{ or } 3.7 \times 10^9 \text{ mm Ag}$$